

= Hints For Naming Ions =

1. All Metals will form cations (+)
2. Group 1A: +1
Group 2A: +2

The name for all these ions will be the same as the element itself.

Examples:

Na⁺ : sodium

Mg²⁺ : magnesium

3. Metals with more than one charge

Example	Stock Naming System (Roman numerals)	Traditional Naming System (Latin root + "ic" or "ous")
Fe ³⁺	iron(III)	ferric (use suffix "-ic" for higher charge)
Fe ²⁺	iron(II)	ferrous (use suffix "-ous" for lower charge)

4. Nonmetals will be anions (-)

Group 5A: -3

Group 6A: -2

Group 7A: -1

To name these ions, change the ending of the element's name to -ide.

Examples:

S²⁻ : sulfide

N³⁻ : nitride

O²⁻ : oxide

5. Oxyanions containing the same nonmetal

The oxyanion with more oxygen atoms: ends with -ate

The oxyanion with fewer oxygens atoms ends with -ite.

Examples:

SO₄²⁻ is sulfate

SO₃²⁻ is sulfite