Copy and complete the following table into your class notebook with the heading: Ionic Formulas

(See text Sections 6.3 & 7.1)

Key

Part 1

For each of the following compounds:

i.) write the name of the compound given
ii.) write the symbol for the ions the compound contains
ii.) name the compound

		Ion Symbols	Chemical Formula
1.)	Example: Lithium carbonate	Li^+ , CO_3^{2-}	Li_2CO_3
2.)	Cuprous sulfite	Cu^+ SO_3^{2-}	Cu_2SO_3
3.)	Cupric sulfite	Cu^{2+} SO_3^{2-}	$CuSO_3$
4.)	Lead(II) nitrite	Pb^{2+} NO_2^-	$Pb(NO_2)_2$
5.)	Aluminum thiosulfate	Al^{3+} $S_2O_3^{2-}$	$Al_2(S_2O_3)_3$
6.)	Ammonium dichromate	NH_4^+ $Cr_2O_7^{2-}$	$(NH_4)_2Cr_2O_7$
7.)	Barium bromide	Ba^{2+} Br^-	$BaBr_2$
8.)	Stannous hydroxide	Sn^{2+} OH^-	$Sn(OH)_2$
9.)	Chromic chromate	Cr^{3+} CrO_4^{2-}	$Cr_2(CrO_4)_3$
10.)	Potassium permanganate		
11.)	Mercurous acetate		
12.)	Magnesium oxalate		
13.)	Cesium bisulfate		
14.)	Zinc perchlorate		
15.)	Hydrogen sulfide		
16.)	Manganous phosphate		
17.)	Calcium hydrogen phosphate		
18.)	Sodium thiosulfate		

Part 2

For each of the following compounds:

i.) copy the formula for the compound given
ii.) write the symbols for the ions in the compound
ii.) name the compound

	Ion Symbols	Compound Name
Example: $SrSO_4$	Sr^{2+} , SO_4^{2-}	Strontium sulfate
19.) $ZnBr_2$	Zn^{2+} Br^-	Zinc bromide
20.) Ag_2S	Ag^+ S^{2-}	Silver sulfide
21.) K_3PO_4	K^+ PO_4^{3-}	Potassium phosphate
22.) $Zn(ClO_3)_2$	Zn^{2+} ClO_3^-	Zinc chlorate
23.) $NH_4C_2H_3O_2$		
24.) $Ca_3(PO_4)_2$		
25.) $Sn(NO_3)_4$		
26.) $MgSO_4$		

Part 3

For each of the following compounds:

i.) copy the formulas for the ions given
ii.) write the chemical formula for the compound that would form
ii.) name the compound

	Chemical Formula	Compound Name
Example: Pb^{4+}, OH^-	$Pb(OH)_4$	Plumbic or Lead(IV) hydroxide
27.) Na^+ , F^-	NaF	Sodium fluoride
28.) K^+ , S^{2-}	K_2S	Potassium sulfide
29.) Ba^{2+} , Br^-	$BaBr_2$	Barium bromide
30.) Al^{3+}, O^{2-}	Al_2O_3	Aluminum oxide
31.) Ca^{2+} , OH^-		
32.) NH_4^+ , P^{3-}		
33.) Cu^+ , NO_3^-		
34.) Cu^{2+} , NO_3^-		
35.) Fe^{3+} , SO_4^{2-}		
36.) Li^+ , CO_3^{2-}		