

# Chemistry

Name: \_\_\_\_\_

Date: \_\_\_\_\_ Block \_\_\_\_\_

## Lab: Double-Replacement Reactions

RXN	Evidence of Reaction
1	
2	
3	
4	
5	
6	
7	
8	
9	
10	

RXN	Balanced Chemical Equations (Include Phases)
1	$\text{NaOH(aq)} + \text{CuSO}_4\text{(aq)} \rightarrow$
2	$\text{NaCl(aq)} + \text{KNO}_3\text{(aq)} \rightarrow$
3	$\text{Na}_2\text{CO}_3\text{(aq)} + \text{HCl(aq)} \rightarrow$
4	$\text{BaCl}_2\text{(aq)} + \text{Na}_2\text{SO}_4\text{(aq)} \rightarrow$
5	$\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow$
6	$\text{Zn(NO}_3)_2\text{(aq)} + \text{CuSO}_4\text{(aq)} \rightarrow$
7	$\text{NaOH(aq)} + \text{FeCl}_3\text{(aq)} \rightarrow$
8	$\text{Na}_2\text{SO}_3\text{(aq)} + \text{HCl(aq)} \rightarrow$
9	$\text{NH}_4\text{Cl(aq)} + \text{CuSO}_4\text{(aq)} \rightarrow$
10	$\text{Pb(NO}_3)_2\text{(aq)} + \text{KI(aq)} \rightarrow$