

# Atoms and Molecules: Chemical Reactions

## What defines a chemical reaction?



### Part 1: Wet Lab.

**A) Safety.** Listen to the safety rules for today's experiment. Write one of them down here:

**B) Observations.** Write down your observations of the three substances:

*Baking soda –*

*Calcium chloride –*

*Phenol red solution –*

### C) Procedure.

Step 1. Put 1 teaspoon of baking soda into a sealable bag.

Step 2. Put 2 teaspoons of calcium chloride into the same bag.

Step 3. Place a test tube with 10 mL of phenol red solution into the bag and hold it upright.

Step 4. Remove the top of the tube. While holding the tube upright, squeeze all the air out of the bag and seal the bag. (The test tube will stay in the bag.) Have your partner make sure the bag is well sealed.

Step 5. Tip the tube of phenol red solution onto the solids. Mix gently from the outside of the bag with your fingertips. (You can let the test tube fall to the bottom.)

**D) Results.** Write down your observations of the reaction:

### E) Further Experiments.

You will now do 2 more experiments of your own design to figure out which reactants are required to produce the heat, gas, or color change.

<b>EXPERIMENT 1</b>		
Circle your choices:	baking soda phenol red solution	calcium chloride water
Write down what happened. Were heat, gas, or a color change produced?		

<b>EXPERIMENT 2</b>		
Circle your choices:	baking soda phenol red solution	calcium chloride water
Write down what happened. Were heat, gas, or a color change produced?		



### F) Class Conclusions from Further Experiments.

Mixing \_\_\_\_\_ produced heat.  
Mixing \_\_\_\_\_ produced a color change.  
Mixing \_\_\_\_\_ produced a gas.

Not all chemical reactions produce heat, a color change, or a gas. The one thing that defines all chemical reactions is that NEW substances are produced.

## The BIG Conclusion: A chemical reaction produces NEW substances!

### G) Reactants and Products.

What did we start with, and what are the new substances that we made today?

<b>Reactants</b> (What we put in the bag)	$H_2O$	<b>Products</b> (What ended up in the bag)
1.	→	1.
2.		2.
		3.
		4.



Matter can change in appearance. Is it a physical change or a chemical change? Here's how to decide:

**4) Physical change** - molecules are the same before and after the change, although the matter may look different.

Examples:

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Hints:

- 1) Physical changes include making mixtures, dissolving one thing in another, and cutting or breaking something.
- 2) All **changes of state** are physical changes. A water molecule is the same water molecule when it is ice, when it is liquid water, and when it is water vapor in the air.

**5) Chemical change** - new and different molecules are formed.

Examples:

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Hints:

- 1) All **chemical reactions** are chemical changes.
- 2) New properties appear.
- 3) The bonds between the atoms are broken and the atoms recombine in new ways.

## B) Modeling a chemical reaction.

Directions:

- 1) Write the formulas for the molecules on the lines below.
- 2) Build and place each LEGO molecule on its formula using the “Chemical Reactants” and “Chemical Products” cards.

$H_2O$   
**CHEMICAL REACTION!**

### Reactants

(What we put in the bag)

\_\_\_\_\_ baking soda molecule

\_\_\_\_\_ baking soda molecule

\_\_\_\_\_ calcium chloride molecule

### Products

(What ended up in the bag)



\_\_\_\_\_ salt molecule

\_\_\_\_\_ salt molecule

\_\_\_\_\_ chalk molecule

\_\_\_\_\_ carbon dioxide molecule

\_\_\_\_\_ water molecule



Look!  
Different molecules have appeared in the bag!

### C) Practice Writing Chemical Formulas.

A chemical formula is an easy way to tell what atoms are present in a compound.

Use the "Atom Key" to find the **chemical symbol** for each element.

It is important to write your formula using the correct uppercase or lowercase letters. The subscript number refers to the atom before it. Remember that "H<sub>2</sub>O" means there are 2 hydrogen atoms and 1 oxygen atom. We write the subscript 2 for the hydrogen but it is unnecessary to write the 1 after the oxygen.

Chemists have a complicated set of rules about the order of atoms in their formulas. For this activity, we'll keep it simple, and list the atoms in order starting from the top of the Atom Key.

#### Directions.

- 1) Watch your teacher demonstrate how to write a formula.
- 2) Build a compound with less than 10 LEGO bricks. (Don't worry about whether it would be a real compound. Build any shape/color you like!)
- 3) Write out the formula for YOUR compound here (write the symbols in the order of the Atom Key, from top to bottom):

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- 4) Trade your compound with your teammate and write out the formula for your TEAMMATE's compound here:

Compare answers with your teammate. Do you agree? Y/N

- 5) Build a second molecule and name it.

\_\_\_\_\_

My formula

\_\_\_\_\_

My Teammate's formula

Look! These formulas follow more complicated rules but are still neat to see!

$\text{CH}_3\text{COOH}$ is the formula for vinegar!	$\text{C}_{19}\text{H}_{14}\text{O}_5\text{S}$ is the formula for phenolsulfonphthalein or phenol red!	$\text{CH}_4$ is the formula for methane gas!	$\text{C}_6\text{H}_{12}\text{O}_6$ is the formula for glucose!	$\text{NaOCl}$ is the formula for bleach!
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