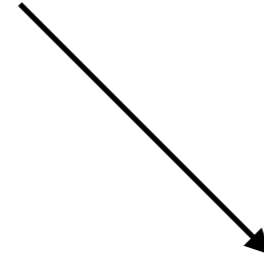
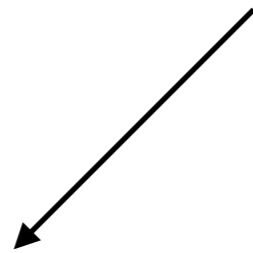


# Naming Acids



## **Acids (produce $H^+$ ions in solution)**



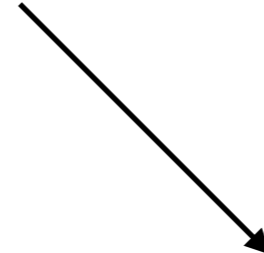
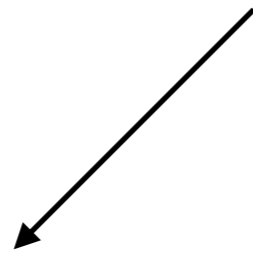
### **Binary Acid:**

$H^+$  and an anion  
(no oxygen)

### **Oxyacid:**

$H^+$  and oxyanion

## Acids (produce $H^+$ ions in solution)



### Binary Acid:

$H^+$  and an anion  
(no oxygen)

### Oxyacid:

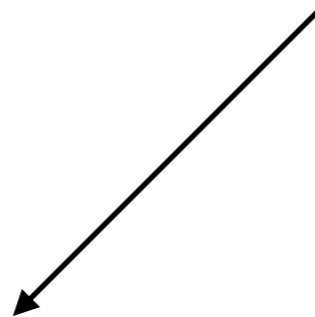
$H^+$  and oxyanion

**hydro** - “root name” - **ic** acid

**HF**

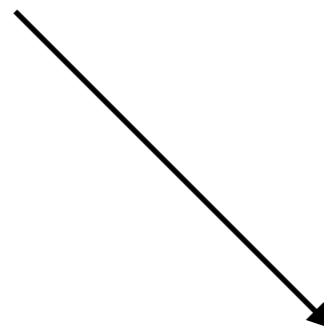
**hydrofluoric** acid

# Acids (produce $H^+$ ions in solution)



## Binary Acid:

$H^+$  and an anion



## Oxyacid:

$H^+$  and oxyanion

change the ending of the oxyanion

-ate  $\longrightarrow$  -ic

-ite  $\longrightarrow$  -ous



nitrate  $\longrightarrow$  nitric

nitric acid

I "ate" something "icky".  
It's "alright" with "us".

## Special Names:

anion	root	example
<b>phosphide/ phosphate</b>	"phosphor"	phosphoric acid
<b>sulfide/sulfate/ sulfite</b>	"sulfur"	sulfuric acid

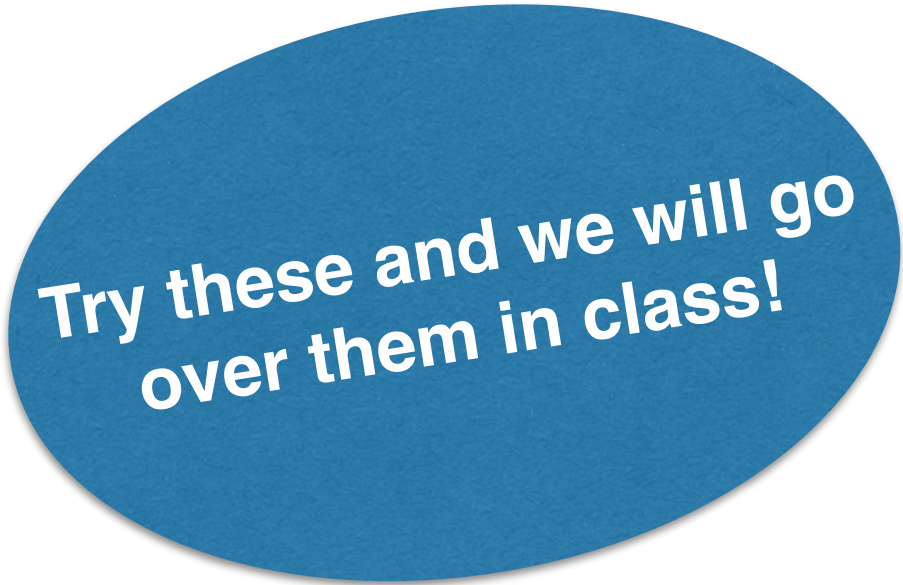
# Practice

1) **HI**

2) **H<sub>3</sub>P**

3) **HClO<sub>2</sub>**

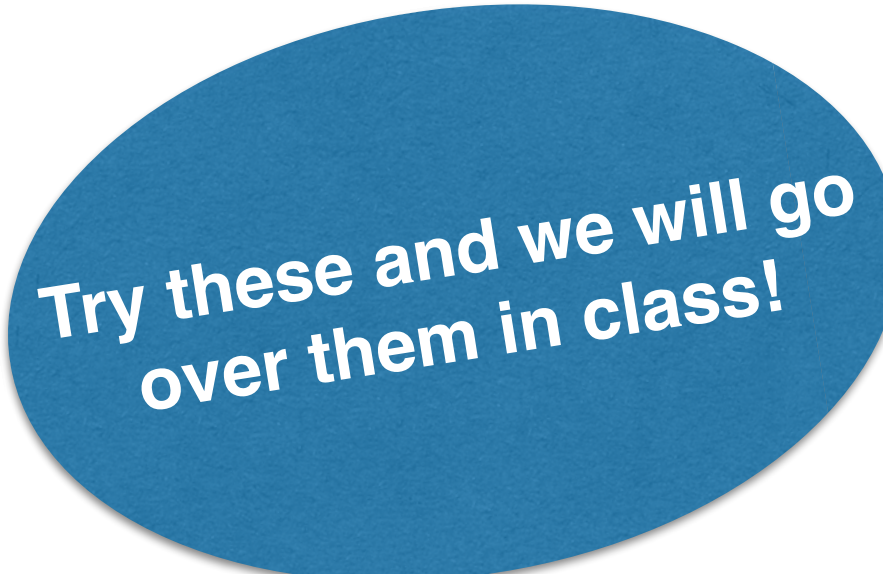
4) **HClO<sub>3</sub>**



Try these and we will go over them in class!

# Practice

- 1) **hydrocyanic acid**
- 2) **hydroselenic acid**
- 3) **perchloric acid**
- 4) **acetic acid**



Try these and we will go over them in class!