Quiz 6.1: Mole Conversions (Practice)

- **Don't forget to head your paper.**
- **For credit, place your answers in the boxes or lines provided. If none, be sure to box your answer.**
- **You must show all work, using proper dimensional analysis, should you want credit for your answers.**
- **Don't forget units and to consider significant figures when writing your answer.**

1. What is another name for the quantity 6.02×10^{23}?

Answer

2. Fill in the empty boxes below with the right names of representative particles & correct molar masses.

Substance	Representative Particle	Molar Mass
Ag		
CO_2		
Li_2O		

3. Complete the following mole conversions, being sure to show all work as taught in class and including units. For credit, box your answer and consider significant figures.

   a. 2.71×10^{24} Formula Units $Ni(ClO_4)_2$ = __?__ g $Ni(ClO_4)_2$

   b.   20.0 g C_2H_6 = _?_ C atoms

   c.   127.0 g Au = _?_ mol Au

   d.   0.500 mol Mg = _?_ atoms Mg