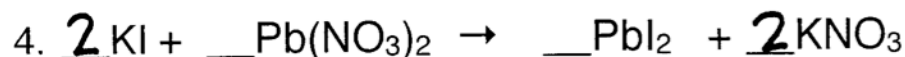
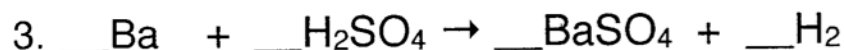
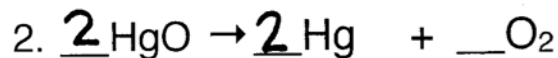
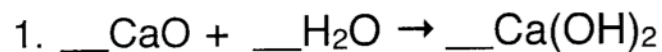


Section 8.1 Copy and complete the following assignment in your class notebook with the heading: Balancing Equations (I)

I. Balance each of the following chemical equations:



5.

6.

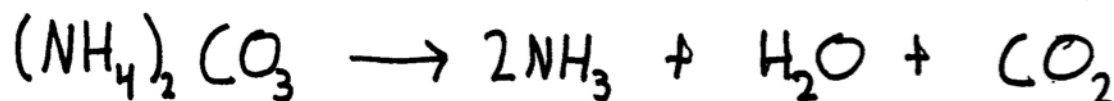
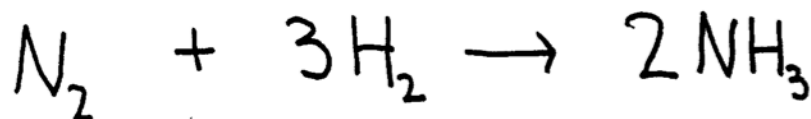
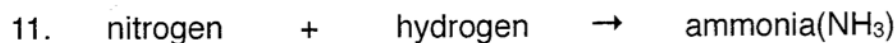
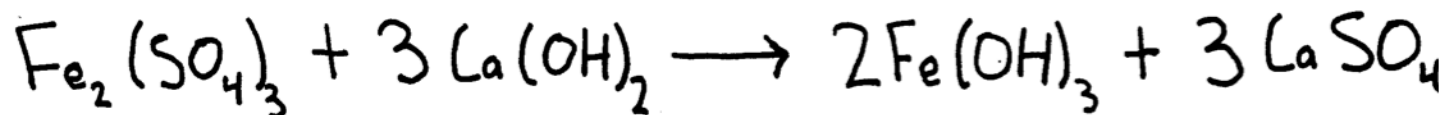
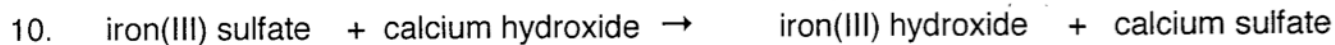
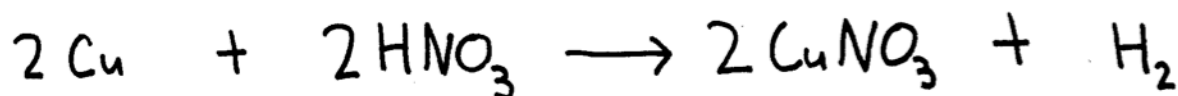
7.

8.

II. a. Copy the word equations given

b. Write the correct formulas for each reactant and product.

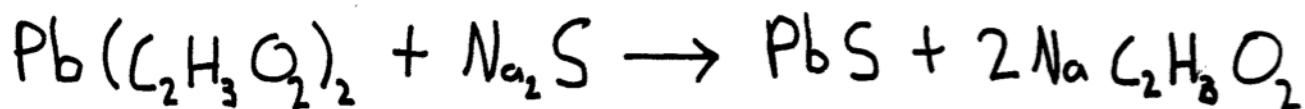
c. Balance the equation.



Key

Blank spaces  
mean "1"

13. lead(II) acetate + sodium sulfide → lead(II) sulfide + sodium acetate



14. cesium + phosphoric acid ( $\text{H}_3\text{PO}_4$ ) → cesium phosphate + hydrogen



15. aluminum oxide → aluminum + oxygen

16. ammonium dichromate + tin(II) sulfate → ammonium sulfate + tin(II) dichromate

17. iron(II) iodide + aluminum phosphate → iron(II) phosphate + aluminum iodide

18. tin + sulfuric acid → tin(II) sulfate + hydrogen

19. propane ( $\text{C}_3\text{H}_8$ ) + oxygen → carbon dioxide + water

20. benzene ( $\text{C}_6\text{H}_6$ ) + oxygen → carbon dioxide + water