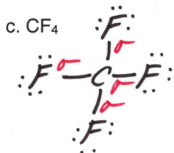
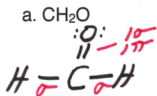


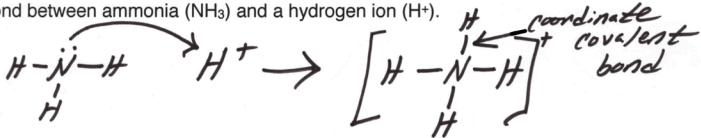
Key

Complete the following in your class notebook with the heading: Covalent compounds (part 2)

1. Draw Lewis structures for the following compounds and label all pi and sigma bonds.



2. Use Lewis structures and arrows to show the formation of a coordinate covalent bond between ammonia (NH_3) and a hydrogen ion (H^+).

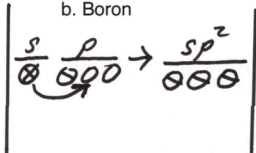


3. Show how each of the following elements form hybridized orbitals (label the orbitals both before and after hybridization takes place):

a. Beryllium

b. Boron

c. Carbon



4. Draw Lewis structures for the following compounds and label the hybridized bonding as sp , sp^2 , or sp^3 .

a. CCl_4

b. BBr_3

c. BeF_2

