

Section 4.3

Complete the following assignment in your class notebook with the heading: **Electron configurations**

Write the electron configuration for each of the following elements and indicate the number of unpaired electrons.

Example: Potassium (atomic #19): $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$; 1 unpaired electron: $\frac{4s}{\text{O}}$

1. Calcium (atomic #20)
2. Lithium (atomic #3)
3. Argon (atomic #18)
4. Iron (atomic #26)
5. Sodium (atomic #11)
6. Sulfur (atomic #16)
7. Iodine (atomic #53)
8. Dysprosium (atomic #66)
9. Radium (atomic #88)
10. Fermium (atomic #100)