

Chemistry

Practice – How Atoms Differ

Name: _____

Block _____ Date Due _____

For each statement below, write true or false.

- _____ The number of neutrons in an atom is referred to as its atomic number.
- _____ The periodic table is arranged by increasing atomic number.
- _____ Atomic number is equal to the number of electrons in a neutral atom.
- _____ The number of protons in an atom identifies it as an atom of a particular element.
- _____ The number of neutrons in an atom is always equal to that atom's atomic number.

Answer the following questions about atoms.

- Lead has an atomic number of 82. How many protons and electrons does lead have? _____
- Oxygen has 8 electrons. How many protons does oxygen have? _____
- Zinc has 30 protons. What is its atomic number? _____
- Astatine has 85 protons. What is its atomic number? _____
- Rutherfordium has an atomic number of 104. How many protons it have? _____
- Polonium has an atomic number of 84. How many electrons does it have? _____
- Niobium has an atomic number of 41. How many neutrons does niobium-93 have? _____


Determine the number of protons, electrons, and neutrons for each isotope described below.

- An isotope with atomic number 19 and mass number 39. $p =$ _____ $n =$ _____ $e^- =$ _____
- An isotope with 14 electrons and a mass number of 28. $p =$ _____ $n =$ _____ $e^- =$ _____
- An isotope with 21 neutrons and a mass number of 40. $p =$ _____ $n =$ _____ $e^- =$ _____
- An isotope with an atomic number of 51 and a mass number of 123. $p =$ _____ $n =$ _____ $e^- =$ _____
- Which isotopes in #13-16 are isotopes of the same element? : 13 14 15 16 Identify the element: _____
circle answer(s)

Write each isotope below in symbolic notation. $\begin{matrix} A \\ Z \\ X \end{matrix}$

18. neon-22 19. helium-5 20. cesium-133 21. uranium-234

Label the mass number and the atomic number on the following symbolic notation.

22. _____  $\begin{matrix} 24 \\ 12 \\ \text{Mg} \end{matrix}$
23. _____

Circle the letter of the choice that best completes the statement.

24. The mass of an electron is
a. smaller than the mass of a proton
b. smaller than the mass of a neutron
c. a tiny fraction of the mass of an atom
d. all of the above
25. One atomic mass unit is
a. 1/12 the mass of a carbon-12 atom.
b. exactly the mass of one proton.
c. 1/16 the mass of an oxygen-16 atom.
d. approximately the mass of 1 proton plus 1 neutron.
26. The atomic mass of an atom is usually not a whole number because it accounts for
a. Only the relative abundance of the atom's isotopes.
b. Only the mass of each of the atom's isotopes.
c. The mass of the atoms electrons.
d. Both the relative abundance and the mass of each of the atom's isotopes.

Use the figures shown to answer the following questions.

Osmium
76
Os
190.2

Niobium
41
Nb
92.906

27. What is the atomic number of osmium? _____
28. What is the chemical symbol for niobium? _____
29. What is the atomic mass of osmium? _____
30. What units is the atomic mass reported in? _____
31. How many electrons does a neutral osmium atom have? _____ A neutral niobium atom ? _____

Calculate the atomic mass of each element described below. Then use the periodic table to identify each element.

32.

Isotope	Mass (amu)	Percent Abundance
^{63}X	62.930	69.17
^{65}X	64.928	30.83

Work:

Name of element: _____

33.

Isotope	Mass (amu)	Percent Abundance
^{35}X	34.969	75.77
^{37}X	36.966	24.23

Work:

Name of element: _____

Answer the following.

34. A carbon atom has a mass number of 12 and an atomic number of 6. How many neutrons does it have?

35. An isotope of mercury has 80 protons and 120 neutrons. What is the mass number of this isotope?

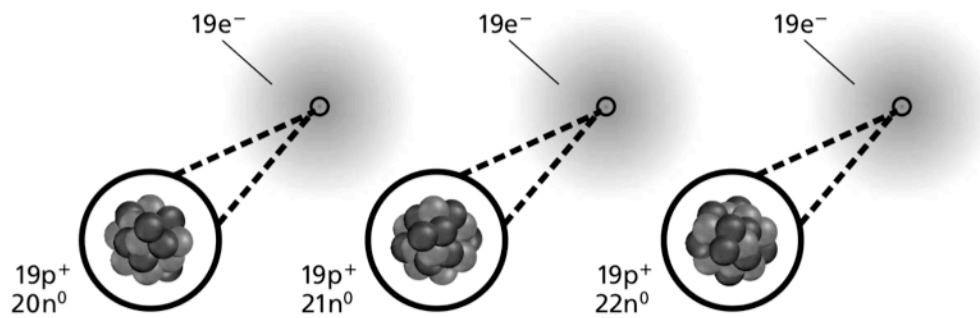
36. An isotope of xenon has an atomic number of 54 and contains 77 neutrons. What is this xenon isotope's mass number?

37. Fill in the following tables, listing how many protons, neutrons, and electrons are contained in each of the following atoms:

Atom	Number of Protons	Number of Neutrons	Number of Electrons
$^{132}_{55}\text{Cs}$			
$^{59}_{27}\text{Co}^{3+}$			
$^{163}_{69}\text{Tm}$			
$^{70}_{30}\text{Zn}^{2+}$			

Atom	Number of Protons	Number of Neutrons	Number of Electrons
Gallium-64			
Fluorine-19			
Titanium-48			
Helium-5			

	Potassium-39	Potassium-40	Potassium-41
Protons	19	19	19
Neutrons	20	21	22
Electrons	19	19	19



38. Which subatomic particles are found in an atom's nucleus?

39. Which subatomic particle identifies an atom as that of a particular element?

40. Explain why atoms are neutral even though they contain charged particles.

41. What do the numbers 39, 40, and 41 after the element name potassium refer to?

42. Write the symbolic notation for each of the following isotopes.

a. potassium-39

b. potassium-40

c. potassium-41

43. Write an equation showing the relationship between an atom's atomic number and its mass number.

44. Lithium has two isotopes: lithium-6 and lithium-7. Draw two diagrams (one for each lithium isotope), like the potassium isotopes at the top of the page.

*Note: Your diagrams should show the protons and neutrons in the nucleus!