

Copy and complete the following table into your class notebook with the heading: Ionic Formulas

(See text Sections 6.3 & 7.1)

Part 1

For each of the following compounds:

- i.) copy the name of the compound given
- ii.) write the symbols for the ions in the compound
- ii.) write the chemical formula for the compound

	<u>Ion Symbols</u>	<u>Chemical Formula</u>
1.) <u>Example</u> : Lithium carbonate	<u>Li⁺ , CO₃²⁻</u>	<u>Li₂CO₃</u>
2.) Copper(I) sulfite	_____	_____
3.) Copper(II) sulfite	_____	_____
4.) Lead(II) nitrite	_____	_____
5.) Aluminum thiosulfate	_____	_____
6.) Ammonium dichromate	_____	_____
7.) Barium bromide	_____	_____
8.) Tin(II) hydroxide	_____	_____
9.) Chromium(III) chromate	_____	_____
10.) Potassium permanganate	_____	_____
11.) Mercury(I) acetate	_____	_____
12.) Magnesium oxalate	_____	_____
13.) Cesium bisulfate	_____	_____
14.) Zinc perchlorate	_____	_____
15.) Hydrogen sulfide	_____	_____
16.) Manganese(II) phosphate	_____	_____
17.) Calcium biphosphate	_____	_____
18.) Sodium thiosulfate	_____	_____

