

Section 6.4 Complete the following table in your class notebook  
with the heading: Molecular Geometry

Valence e <sup>-</sup> Total	Lewis Dot Structure	Molecular Shape & Bond Angle	Name of Shape	Molecular Polarity
PCl <sub>3</sub>				
H <sub>2</sub> S				
CBr <sub>4</sub>				
BrCN				
SO <sub>3</sub> <sup>2-</sup>				
C <sub>2</sub> F <sub>2</sub>			(with respect to either C)	
C <sub>2</sub> H <sub>4</sub>			(with respect to either C)	

Valence e <sup>-</sup> Total	Lewis Dot Structure	Molecular Shape & Bond Angle	Name of Shape	Molecular Polarity
BF <sub>3</sub>	(an exception to the octet rule!)			
OF <sub>2</sub>				
O <sub>3</sub>				
NO <sup>+</sup>				
PO <sub>3</sub> <sup>-</sup>				
H <sub>2</sub> SO <sub>4</sub>			(with respect to S)	
			(with respect to O w/H)	
H <sub>2</sub> CO <sub>3</sub>			(with respect to C)	
			(with respect to O w/H)	
HC <sub>2</sub> H <sub>3</sub> O <sub>2</sub>	(Hint: 3 H are bonded to 1 C)		(with respect to C w/3 H)	
			(with respect to other C)	