

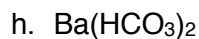
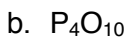
Name:

## WP Practice

### Exam 3: Ionic and Covalent Compounds

(Also review unit 3 and unit 4 pretest packets and naming assignments)

1. Write the name (IUPAC rules) for the following compounds:



2. Write the formula for the following:

a. magnesium sulfate

e. mercury(II) sulfide

b. sodium selenide

f. tin(IV) nitrite

c. dinitrogen monoxide

g. calcium hydroxide

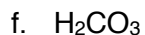
d. sulfur hexafluoride

h. tetraphosphorus trisulfide

3. Write the name or the formula for the following acids:



e. chloric acid



c. sulfuric acid

g. hydrosulfuric acid



h. acetic acid

4. The term which best describes a substance made of cations and delocalized electrons is

- A) covalent compound
- B) molecule
- C) ionic solid
- D) metal
- E) cation

5. The term which best describes a substance made of metal cations and nonmetal anions is

- A) covalent compound
- B) molecule
- C) ionic solid
- D) metal
- E) cation

6. A double bond is a \_\_\_\_\_ bond that involves the sharing of \_\_\_\_\_ electrons.

- A) covalent; two
- B) covalent; four
- C) ionic; two
- D) ionic; six
- E) covalent; six

7. Define the following terms:

a. diatomic elements:

b. molecule:

c. covalent bond:

d. electronegativity:

e. polar covalent:

f. nonpolar covalent:

g. coordinate covalent bond: