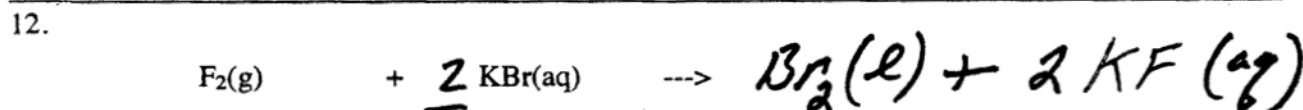
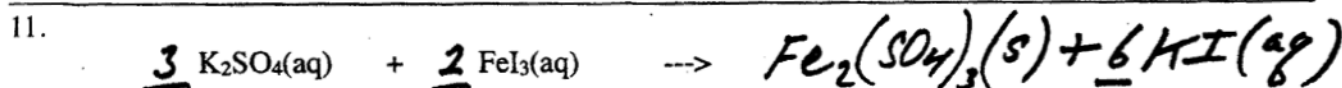
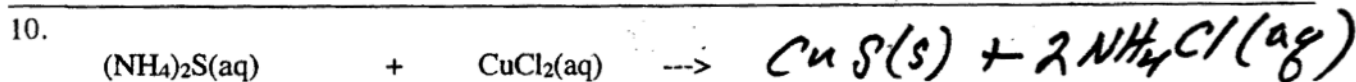
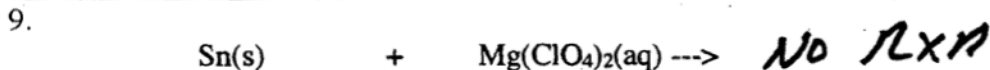
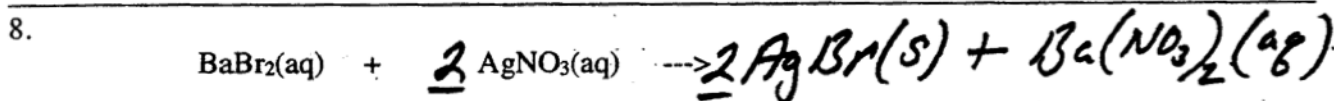
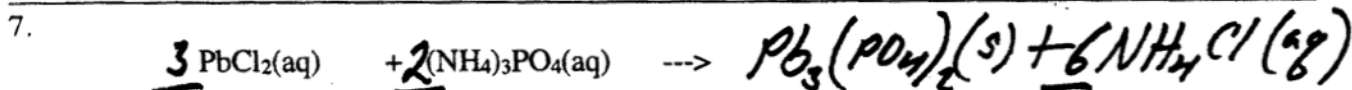
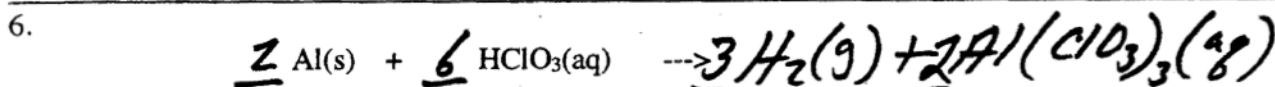
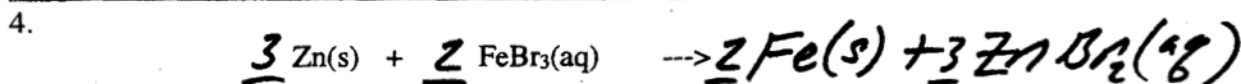
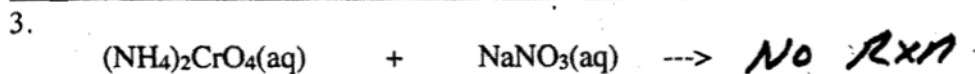
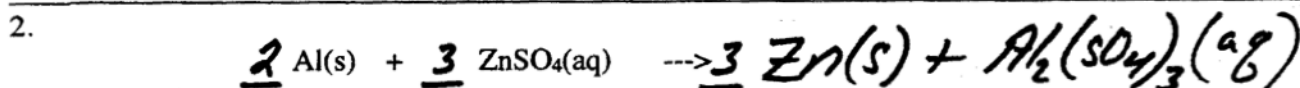
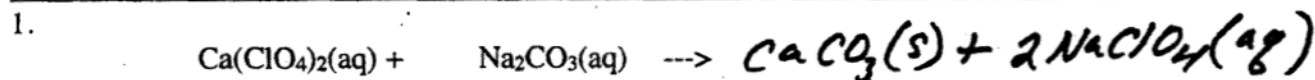


Name \_\_\_\_\_ Block \_\_\_\_\_

## Chemical Equations (IV) =Predicting Products=

*Key*

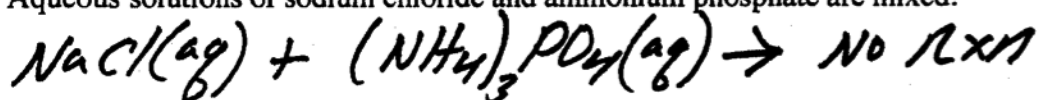
Predict if the following single-replacement or double-replacement reactions will occur. If no reaction takes place, write "No Rxn" and move on to the next problem. If a reaction occurs, write a balanced chemical equation for the reaction, including phases.



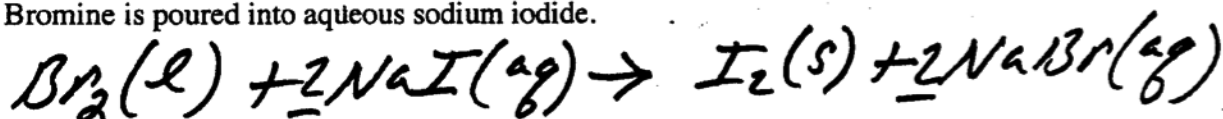
(over)

Write the chemical formulas and phases for reactants in each of the following problems, then determine if the reaction takes place...

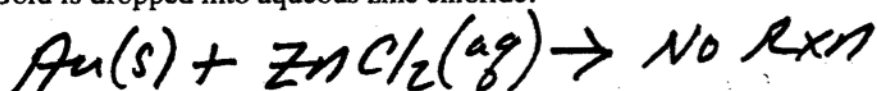
13. Aqueous solutions of sodium chloride and ammonium phosphate are mixed.



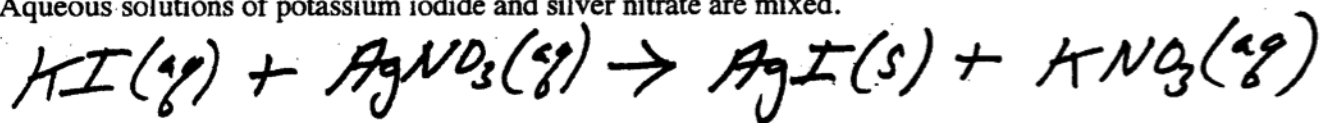
14. Bromine is poured into aqueous sodium iodide.



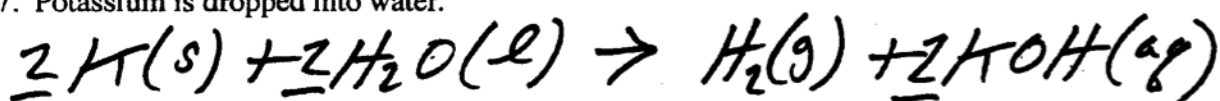
15. Gold is dropped into aqueous zinc chloride.



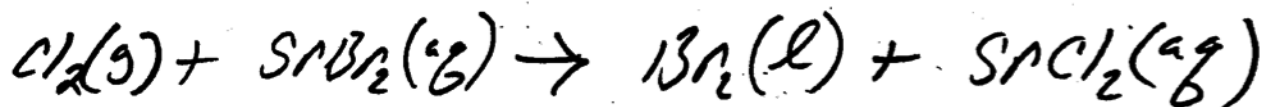
16. Aqueous solutions of potassium iodide and silver nitrate are mixed.



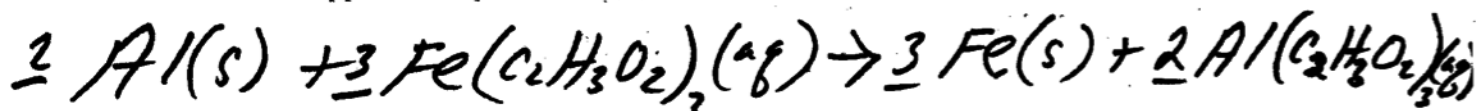
17. Potassium is dropped into water.



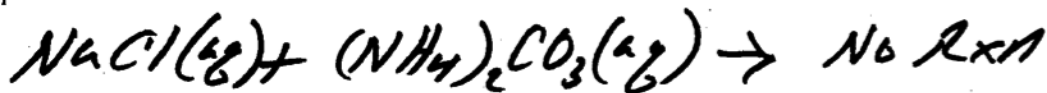
18. Chlorine gas is bubbled into aqueous strontium bromide.



19. Aluminum metal is dropped into aqueous iron(II) acetate.



20. Aqueous solutions of sodium chloride and ammonium carbonate are mixed.



21. Zinc metal is dropped into hydrochloric acid.

