Fill in the table:

	Number of Valence Electrons	Lewis Dot Structure	Number of electrons still <u>needed</u> for stable octet
Fluorine			
Carbon			
Argon			
Hydrogen			
Nitrogen			

So what happens when F and C meet?

Remember….

1) *When a <u>metal</u> meets a <u>nonmetal</u> the metal transfers electrons to the nonmetal, then the resulting ions attract (**ionic bond**).*
2) *When a <u>metal</u> meets another <u>metal</u> they donate their valence electrons to the "sea of electrons" and the resulting metallic cations are attracted to the negative sea of electrons (**metallic bond**).*
3) *But what about <u>2 or more nonmetals</u> meeting? How do they get their octets?*